

**KEY**

## Review: Naming Ionic Compounds

Write the formulas of the following ionic compounds (1 pt each):

- 1) iron (II) arsenide \_\_\_\_\_
- 2) lead (II) sulfate \_\_\_\_\_
- 3) lead (IV) hydroxide \_\_\_\_\_ <sup>C</sup>  
*OH*
- 4) copper (II) acetate \_\_\_\_\_
- 5) beryllium chloride \_\_\_\_\_
- 6) ammonium chromate \_\_\_\_\_
- 7) silver oxide \_\_\_\_\_
- 8) potassium sulfide \_\_\_\_\_

Write the names of the following ionic compounds (1 pt each):

- 9) KI \_\_\_\_\_
- 10)  $Mn_2(SO_3)_7$  \_\_\_\_\_
- 11)  $SnBr_4$  \_\_\_\_\_
- 12)  $Mg_3P_2$  \_\_\_\_\_
- 13) NaF \_\_\_\_\_
- 14)  $Sr(MnO_4)_2$  \_\_\_\_\_
- 15)  $Cr(PO_4)_2$  \_\_\_\_\_
- 16)  $Al_2Se_3$  \_\_\_\_\_

## Review: Naming Ionic Compounds

Write the formulas of the following ionic compounds (1 pt each)

- 1) iron (II) arsenide  $Fe_3As_2$
- 2) lead (II) sulfate  $PbSO_4$
- 3) lead (IV) hydroxide  $Pb(OH)_4$
- 4) copper (II) acetate  $Cu(C_2H_3O_2)_2$
- 5) beryllium chloride  $BeCl_2$
- 6) ammonium chromate  $(NH_4)_2CrO_4$
- 7) silver oxide  $Ag_2O$
- 8) potassium sulfide  $K_2S$

Write the names of the following ionic compounds (1 pt each)

- 9) KI potassium iodide
- 10)  $Mn_2(SO_3)_7$  manganese (VII) sulfite
- 11)  $SnBr_4$  tin (IV) bromide
- 12)  $Mg_3P_2$  magnesium phosphide
- 13) NaF sodium fluoride
- 14)  $Sr(MnO_4)_2$  strontium permanganate
- 15)  $Cr(PO_4)_2$  chromium (II) phosphate
- 16)  $Al_2Se_3$  aluminum selenide

