

$$0.03\text{L} \times \frac{0.1\text{mol AgNO}_3}{1\text{L}} = 0.003\text{mol AgNO}_3$$

$$0.01\text{L} \times \frac{0.1\text{mol MgCl}_2}{1\text{L}} = 0.001\text{mol MgCl}_2$$

LR. \nearrow

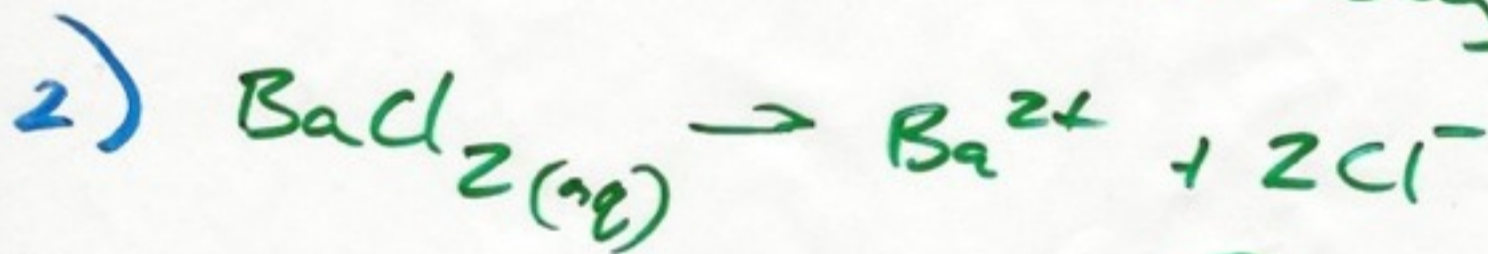
$$0.001\text{mol MgCl}_2 \times \frac{2\text{mol AgNO}_3}{1\text{mol MgCl}_2} =$$

$$0.001\text{mol MgCl}_2 \times \frac{2\text{mol AgCl}}{1\text{mol MgCl}_2} \times \frac{143.3\text{g}}{1\text{mol}}$$

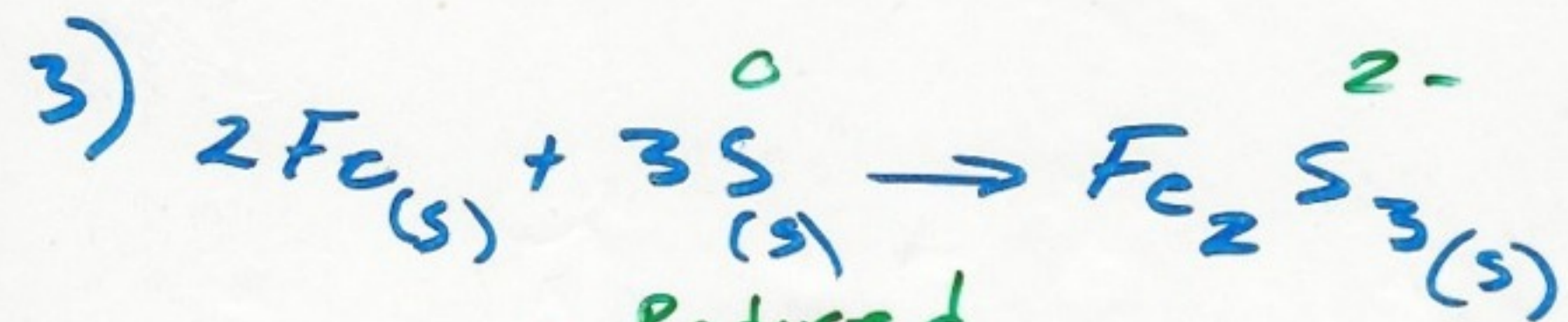
0.002mol
AgNO₃
needed

0

0.286g AgCl



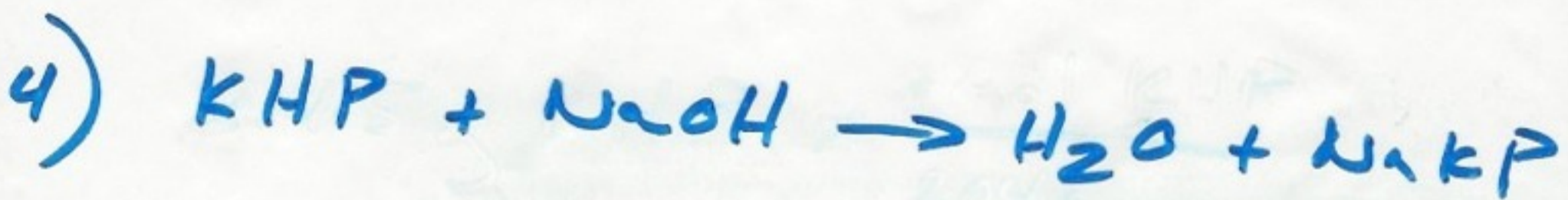
0



- Reduced

- oxidizing Agent

(A)

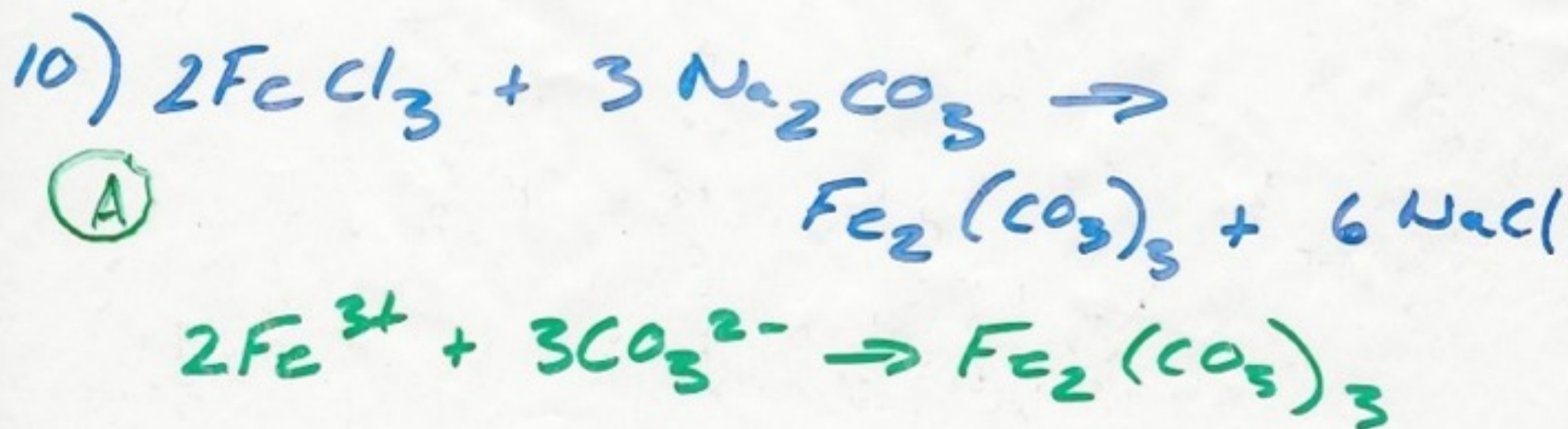
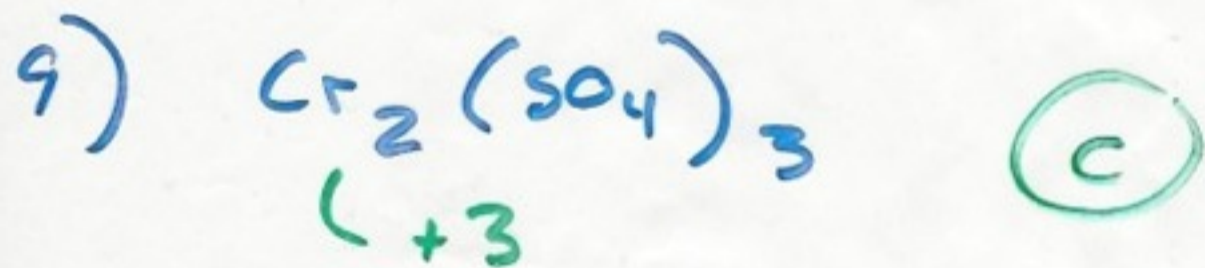
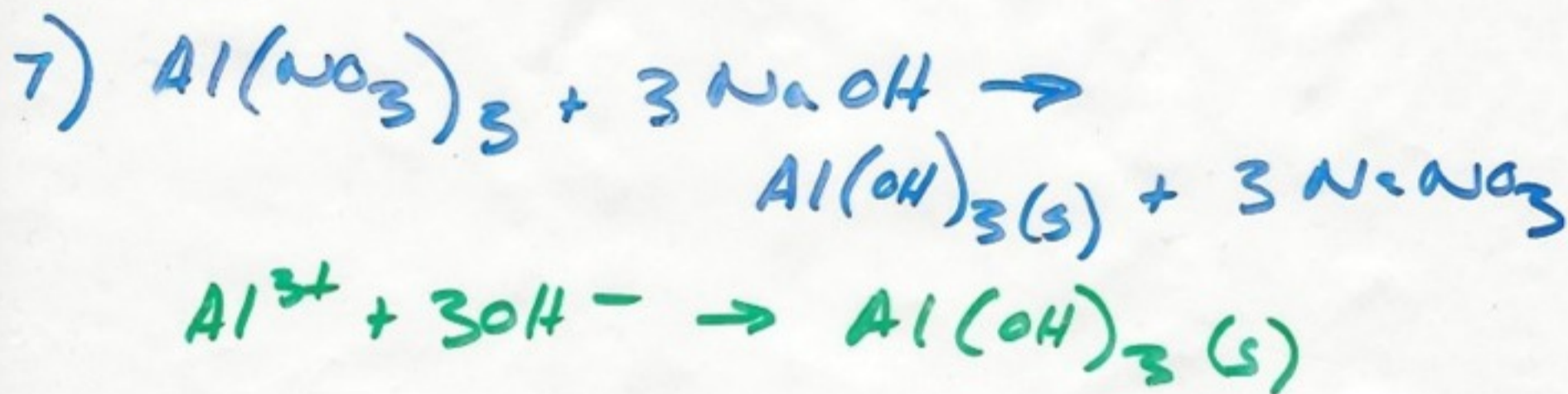
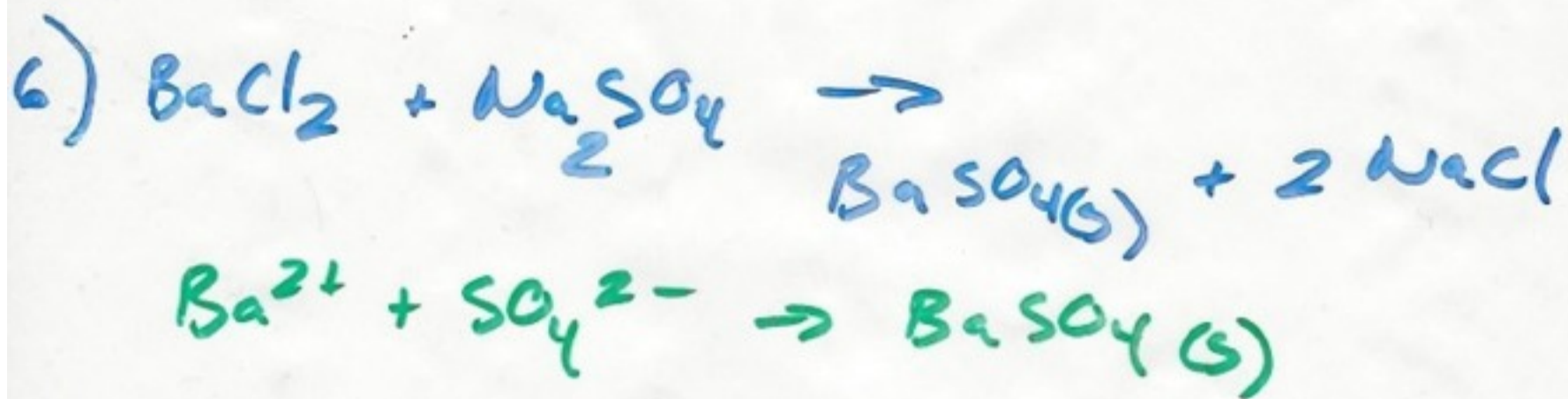
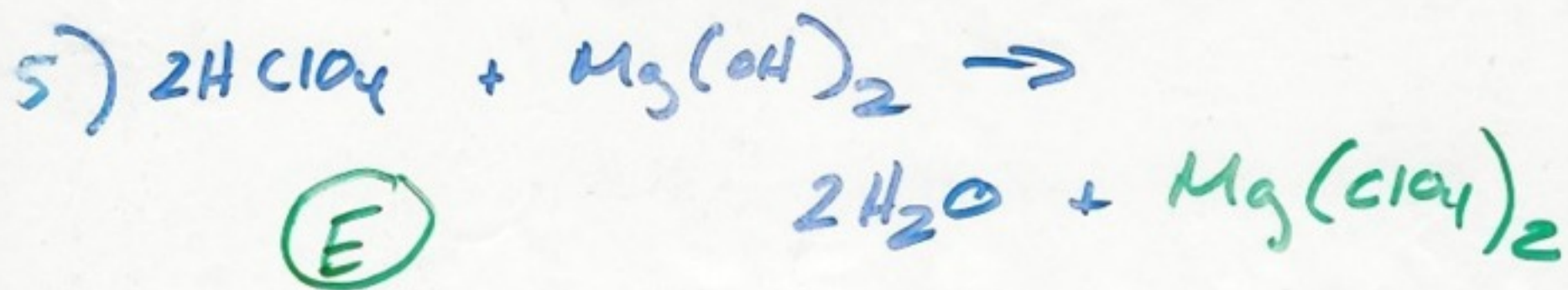


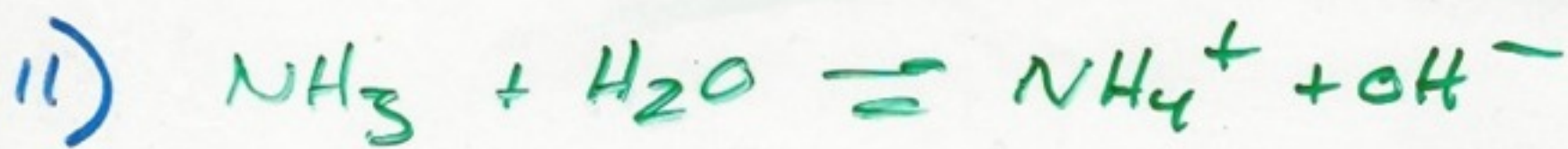
$$2.415\text{g KHP} \times \frac{1\text{mol KHP}}{204\text{g}} \times \frac{1\text{mol NaOH}}{1\text{mol KHP}}$$

$$\times \frac{1}{0.0338\text{L}} = \frac{0.349\text{mol NaOH}}{\text{L}}$$

←

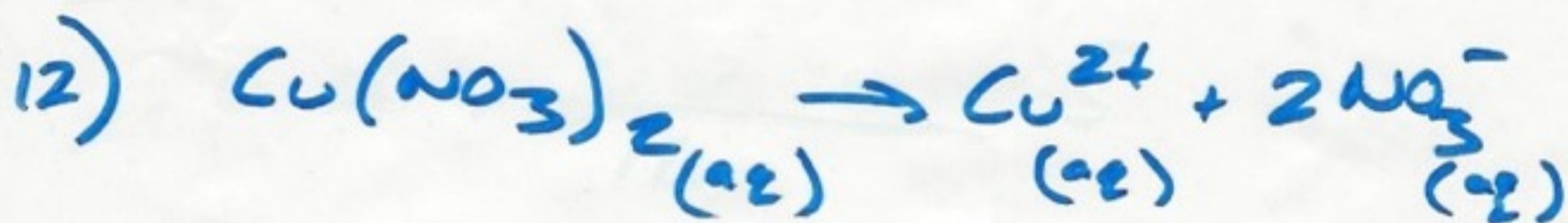
(D)





weak Base
weak electrolyte

(D)



$$0.2\text{L} \times \frac{0.15\text{ mol NO}_3^-}{1\text{L}} \times \frac{1\text{ mol Cu}(\text{NO}_3)_2}{2\text{ mol NO}_3}$$

$$\times \frac{187.5\text{g}}{1\text{ mol Cu}(\text{NO}_3)_2} = 2.8\text{g Cu}(\text{NO}_3)_2$$

(B)