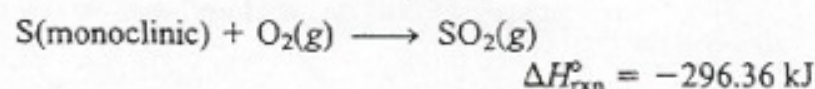
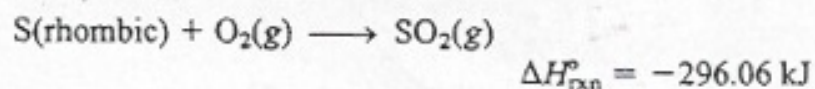


1. _____ From these data,

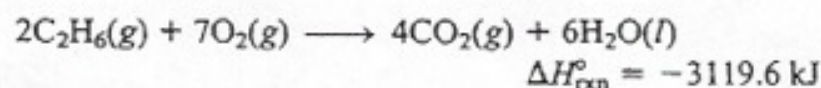
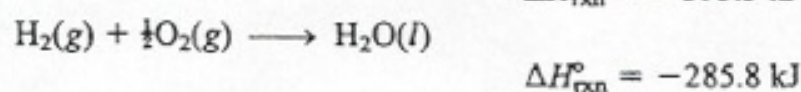
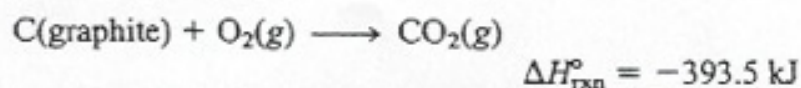


calculate the enthalpy change for transformation

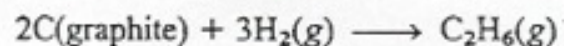


(Monoclinic and rhombic are different allotropic forms of elemental sulfur.)

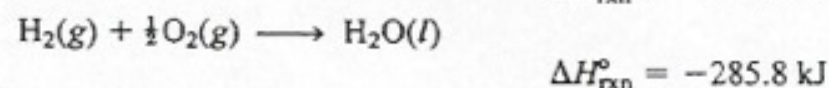
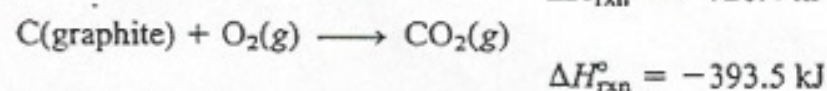
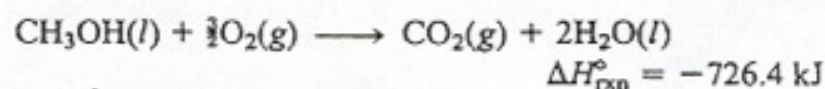
2. _____ From the following data,



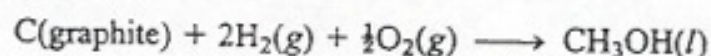
calculate the enthalpy change for the reaction



3. _____ From the following heats of combustion,



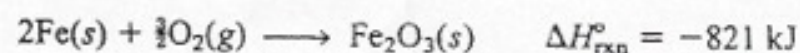
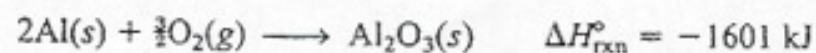
calculate the enthalpy of formation of methanol (CH_3OH) from its elements:

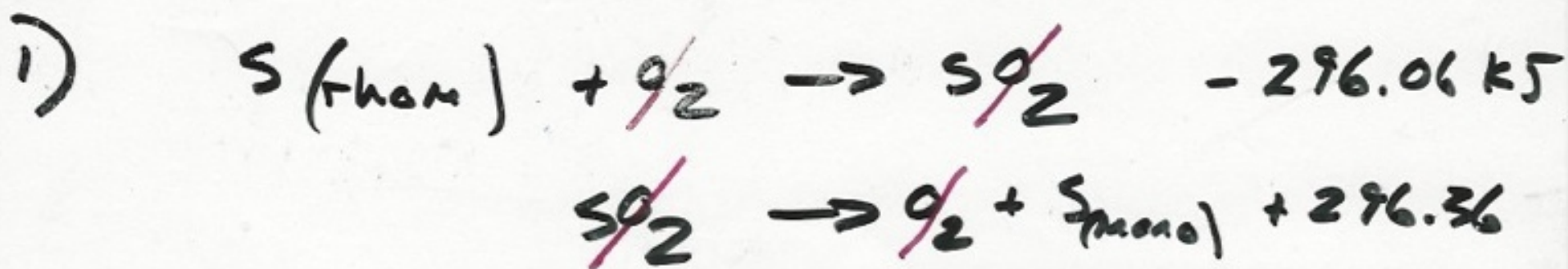


4. _____ Calculate the standard enthalpy change for the reaction

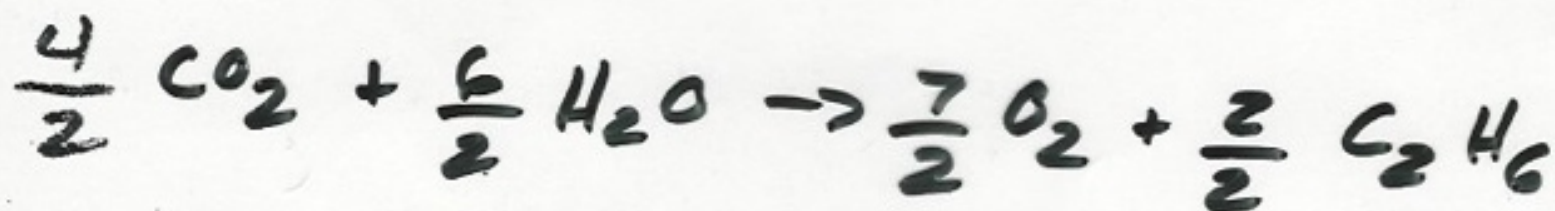
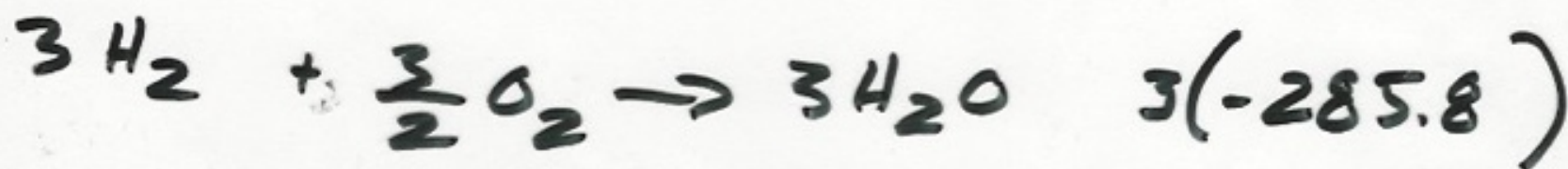


given that





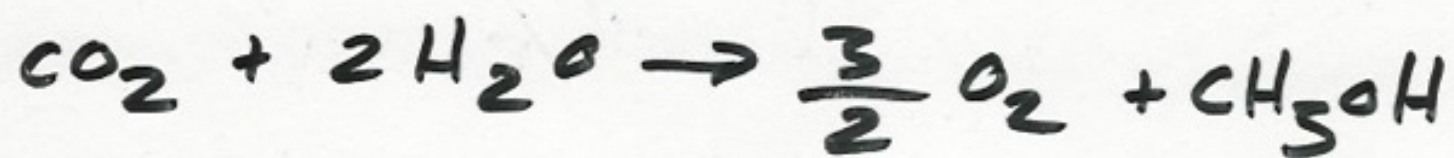
2)



$$\frac{+3119.6}{2}$$



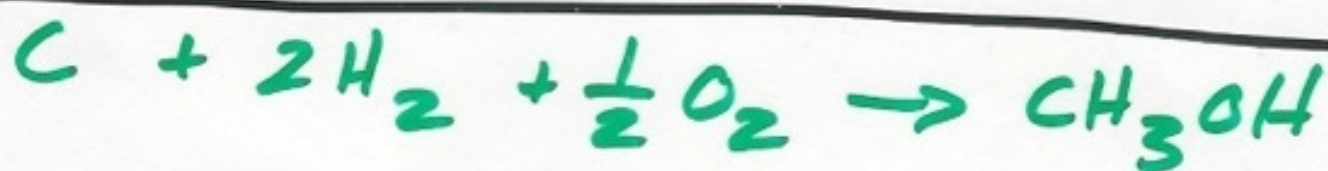
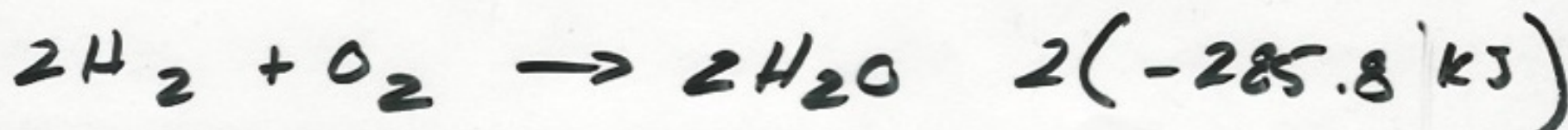
3)



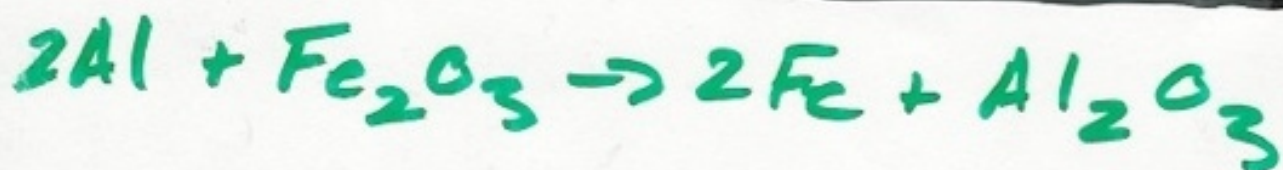
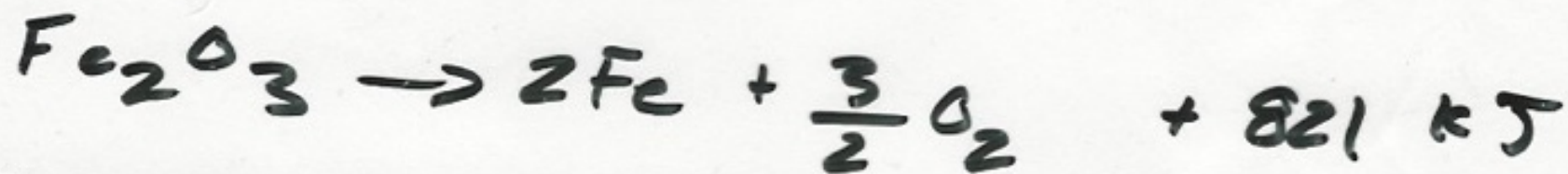
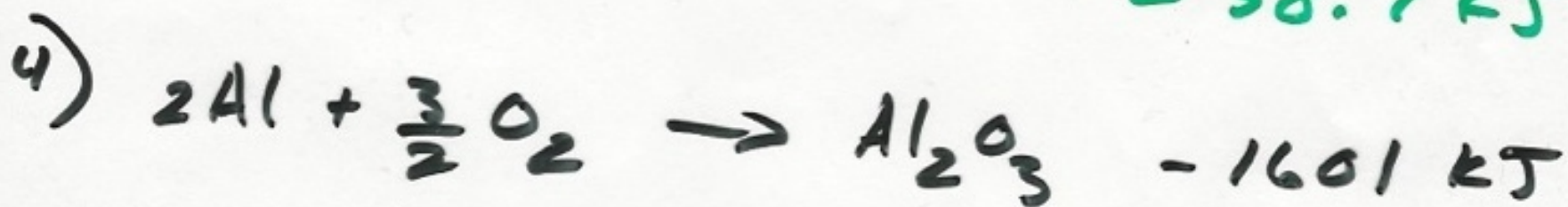
+ 726.4 kJ



- 393.5 kJ



- 238.7 kJ



- 780 kJ

